414 *J. Chem. Eng. Data* **1997,** *42,* 414

Corrections

Vapor–Liquid Equilibria for Acetone + Chloroform + Methanol and Constituent Binary Systems at 101.3 kPa. Toshihiko Hiaki, Kiyofumi Kurihara, and Kazuo Kojima, *J. Chem. Eng. Data* **1994**, *39*, 714–719.

In Table 5 on p 716, the experimental value of x_1 for the acetone (1) + chloroform (2) + methanol (3) system needs correction. The correct value is 0.318 (**not** 0.381).

Line 5 of eq 13 on p 718 should read

$$[-x_3b_{23} + (x_3{}^2 - 2x_2x_3)c_{23} + (x_2 - x_3)(x_3{}^2 - 3x_2x_3)d_{23} +$$

Line 4 of eq 14 on p 718 should read

$$x_3{}^2 - x_1{}^2)e_{31} + ...] + [-x_2b_{12} + (x_2{}^2 - 2x_1x_2)c_{12} +$$

Line 9 of eq 14 on p 718 should read

$$x_1x_2x_3\}c_1 + \{x_1x_2(x_3 - x_1) - x_2x_3(x_3 - x_1) + 2x_1x_2x_3\}c_2 +$$

JE960366K

S0021-9568(96)00366-4

Isopiestic Investigation of the Osmotic and Activity Coefficients of Aqueous NaBr and the Solubility of NaBr·2H$_2$O(cr) at 298.15 K: Thermodynamic Properties of the NaBr + H$_2$O System over Wide Ranges of Temperature and Pressure. Joseph A. Rard and Donald G. Archer, *J. Chem. Eng. Data* **1995**, *40*, 170–185

Table 9 contains smoothed values of the standard state thermodynamic properties of NaBr(aq) at selected temperatures and pressures. However, the given values of the apparent molar heat capacity $C_{p,\phi}$ and the apparent molar volume V_ϕ in this table are actually for 0.1 mol·kg^{-1} and not infinite dilution.

JE9605001

S0021-9568(96)00500-6

Vapor–Liquid Equilibria for Alcohol + Alcohol + Sodium Iodide at 298.15 K. Hideki Yamamoto, Koji Fukase, and Junji Shibata, *J. Chem. Eng. Data* **1996**, *40*, 1066–1070.

Parts of Tables 4 and 5 were identical. Table 4 is correct as given below.

Table 4. Vapor–Liquid Equilibria of Methanol (1) + Propan-1-ol (2) + NaI (3) at 298.15 K

x_1°	x_3	y_1	100δ	P/kPa
0.042	0.000	0.196	−2.96	3.41
0.153	0.000	0.535	4.33	5.10
0.304	0.000	0.731	1.44	7.25
0.351	0.000	0.778	2.05	7.87
0.434	0.000	0.824	0.37	9.16
0.541	0.000	0.886	1.03	10.65
0.605	0.000	0.908	0.53	11.62
0.723	0.000	0.947	0.57	13.21
0.878	0.000	0.980	0.16	15.42
			$100\Delta = \pm 1.46$	

x_1°	x_3	y_1	100δ	P/kPa	γ_1	γ_2
0.084	0.143	0.301	−2.55	3.23	0.765	1.021
0.171	0.129	0.493	0.49	4.23	0.813	1.032
0.332	0.132	0.704	1.03	5.85	0.850	1.035
0.485	0.119	0.817	1.32	7.57	0.892	1.054
0.583	0.116	0.869	0.97	8.71	0.906	1.067
0.661	0.118	0.898	1.15	9.63	0.908	1.079
0.788	0.113	0.945	0.37	11.63	0.921	1.123
0.867	0.107	0.968	0.11	12.55	0.933	1.170
0.920	0.091	0.984	−0.20	13.75	0.958	1.215
			$100\Delta = \pm 0.91$			

[a] x_1° = mole fraction of salt-free base. γ is the activity coefficient using Hála's equation. $\delta = \{(y_{1(\text{cal*})} - y_{1(\text{exp})}\}/y_{1(\text{exp})}$. $\Delta = (1/N)\sum_{i=1}^{N}|\{(y_{1(\text{cal*})} - y_{1(\text{exp})}\}/y_{1(\text{exp})}|_i$. Asterisks refer to smoothed values from experimental values using Hála's equation ($\beta = 3.1$, $E_{13} = 37.18$, $\Lambda_{13} = 0.7882$, $E_{23} = 15.25$, $\Lambda_{23} = 0.1431$, $\Lambda_{12} = 1.8698$, $\Lambda_{21} = 0.3611$).

JE9603640

S0021-9568(96)00364-0

S0021-9568(96)00366-4 CCC: $14.00